ACIDS, BASES & SALTS WORKSHEET - 1

1. Zinc sulphate is called a "salt", sulphur dioxide an "acidic oxide" and lead monoxide a "basic oxide". What is meant by these terms?
2. Give equation and indicate briefly, how will you prepare crystals of zinc sulphate from zinc filings?
3. Name three classes of substances, which react with an acid to form salts. Write equations to describe their reactions with suitable acids.
4. Making use only of substances chosen from those given below:
   Give the equations for the reactions by which you could obtain:
   (i) Hydrogen (ii) Sulphur dioxide
   (iii) Carbon dioxide (iv) Zinc carbonate
5. (i) Name four soluble salts.
   (ii) Name four insoluble salts.
6. Draw the structure of the stable positive ion formed when an acid dissolves in water.
7. Hydrochloric acid is considered as a strong acid whereas acetic acid is a weak acid. Why?
8. Why dilute sulphuric acid is stronger acid than concentrated sulphuric acid?
9. (i) Acetic acid is monobasic. Why?
   (ii) Carbonic acid is a dibasic acid. Why?
   (iii) Sodium hydroxide is a monoacidic base. Why?
10. An aqueous solution of the salt ammonium chloride is acidic in nature while an aqueous solution of sodium chloride is neutral. Why?
11. (i) An aqueous solution of the zinc sulphate acidic in nature. Why?
    (ii) An aqueous solution of ammonium acetate, neutral in nature. Why?
    (iii) An aqueous solution of sodium carbonate is alkaline and that of ammonium chloride is acidic in behavior. Why?