



**SpeedLabs**  
**Science**

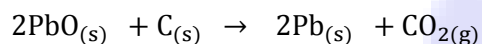
**CBSE 10<sup>th</sup>**

**TEEVRA EDUTECH PVT. LTD.**

# Chemical Reactions and Equations

## Exercise

1. Which of the statements about the reaction below are incorrect?



- (a) Lead is getting reduced.
- (b) Carbon dioxide is getting oxidised.
- (c) Carbon is getting oxidised.
- (d) Lead oxide is getting reduced.

(i) (a) and (b)

(ii) (a) and (c)

(iii) (a), (b) and (c)

(iv) all

**Ans.** (i) (a) and (b)

2.  $\text{Fe}_2\text{O}_3 + 2\text{Al} \rightarrow \text{Al}_2\text{O}_3 + 2\text{Fe}$ . The above reaction is an example of a

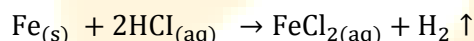
- (a) Combination Reaction.
- (b) Double Displacement Reaction.
- (c) Decomposition Reaction.
- (d) Displacement Reaction.

**Ans.** (d) The given reaction is an example of a displacement reaction.

3. What happens when dilute hydrochloric acid is added to iron filings? Tick the correct answer.

- (a) Hydrogen gas and iron chloride are produced.
- (b) Chlorine gas and iron hydroxide are produced.
- (c) No reaction takes place.
- (d) Iron salt and water are produced.

**Ans.** (a) Hydrogen gas and iron chloride are produced. The reaction is as follows.

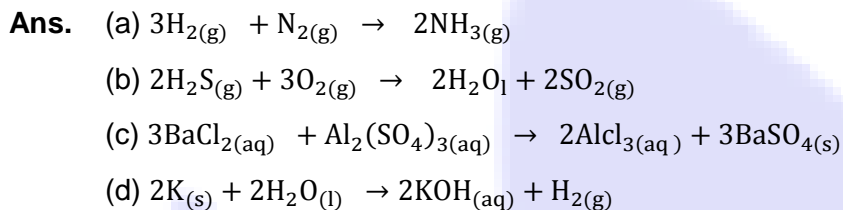


4. what is balanced chemical equation? Why should chemical equations be balanced?

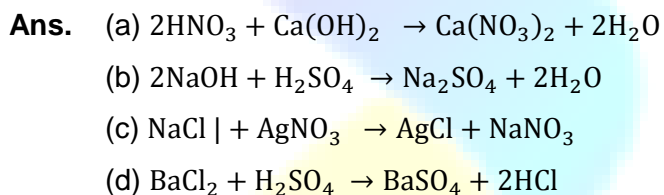
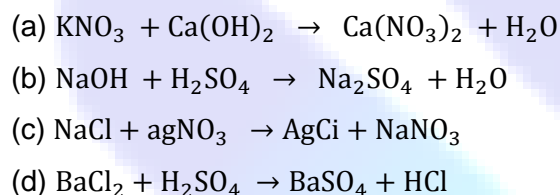
**Ans.** A reaction which has an equal number of atoms of all the elements on both sides of the chemical equation is called a balanced chemical equation.

The law of conservation of mass states that mass can neither be created nor destroyed. Hence, in a chemical reaction, the total mass of reactants should be equal to the total mass of the products. It means that the total number of atoms of each element should be equal on both sides of a chemical equation. Hence, it is for this reason that chemical equations should be balanced.

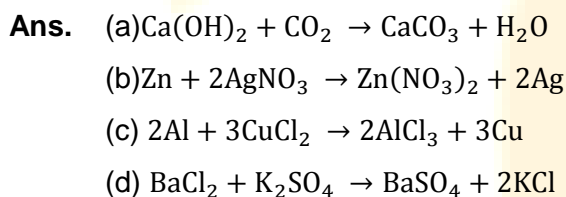
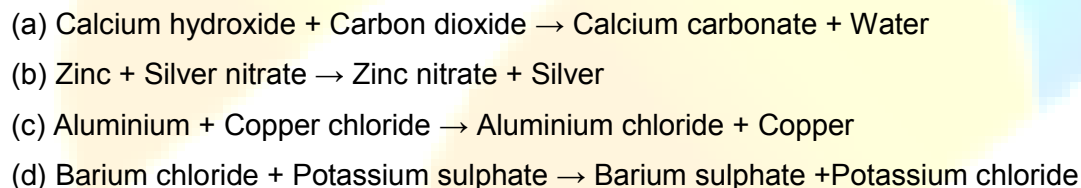
5. Translate the following statements into chemical equations and then balance them.
- Hydrogen gas combines with nitrogen to form ammonia.
  - Hydrogen sulphide gas burns in air to give water and sulphur dioxide.
  - Barium chloride reacts with aluminium sulphate to give aluminium chloride and a precipitate of barium sulphate.
  - Potassium metal reacts with water to give potassium hydroxide and hydrogen gas.



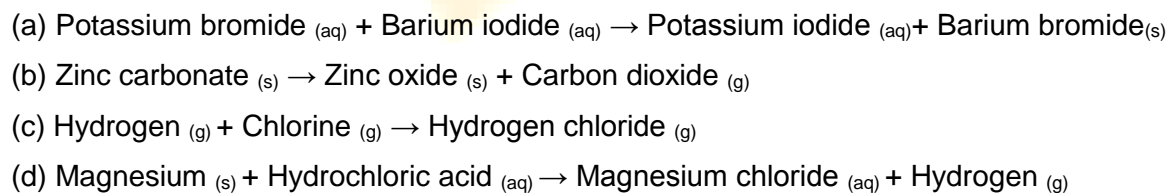
6. Balance the following chemical equations.



7. Write the balanced chemical equations for the following reactions.



8. Write the balanced chemical equation for the following and identify the type of reaction in each case.



**Ans.** (a)  $2\text{KBr}_{(\text{aq})} + \text{BaI}_{2(\text{aq})} \rightarrow 2\text{KI}_{(\text{aq})} + \text{BaBr}_{2(\text{s})}$ ; Double displacement reaction.

(b)  $\text{ZnCO}_{3(\text{s})} \rightarrow \text{ZnO}_{(\text{s})} + \text{CO}_{2}$ ; Decomposition reaction

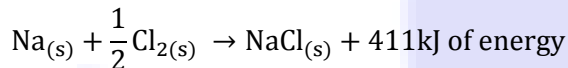
(c)  $\text{H}_{2(\text{g})} + \text{Cl}_{2(\text{g})} \rightarrow 2\text{HCl}_{(\text{g})}$ ; Combination reaction.

(d)  $\text{Mg}_{(\text{s})} + 2\text{HCl}_{(\text{aq})} \rightarrow \text{MgCl}_{2(\text{aq})} + \text{H}_{2(\text{g})}$ ; Displacement reaction

**9.** What does one mean by exothermic and endothermic reactions? Give examples.

**Ans.** Chemical reactions that release energy in the form of heat, light, or sound are called exothermic reactions.

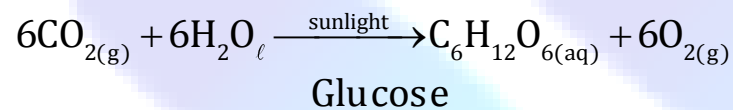
Example. Mixture of sodium and chlorine to yield table salt



In other words, combination reactions are exothermic.

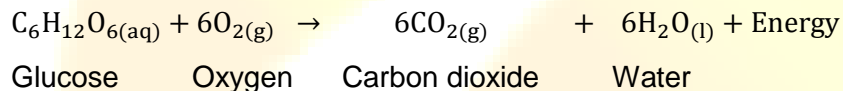
Reactions that absorb energy or require energy in order to proceed are called endothermic reactions.

For example. In the process of photosynthesis, plants use the energy from the sun to convert carbon dioxide and water to glucose and oxygen.



**10.** Why is respiration considered an exothermic reaction? Explain.

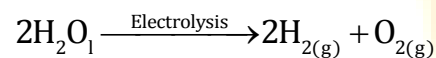
**Ans.** Energy is required to support life. Energy in our body is obtained from the food we eat. During digestion, large molecules of food are broken down into simpler substances such as glucose. Glucose combines with oxygen in the cells and provides energy. The special name of this combustion reaction is respiration. Since energy is released in the whole process, it is an exothermic process.



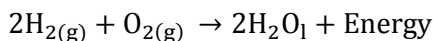
**11.** Why are decomposition reactions called the opposite of combination reactions? Write equations for these reactions.

**Ans.** Decomposition reactions are those in which a compound breaks down to form two or more substances. These reactions require a source of energy to proceed. Thus, they are the exact opposite of combination reactions in which two or more substances combine to give a new substance with the release of energy.

Decomposition reaction.  $\text{AB} + \text{Energy} \rightarrow \text{A} + \text{B}$

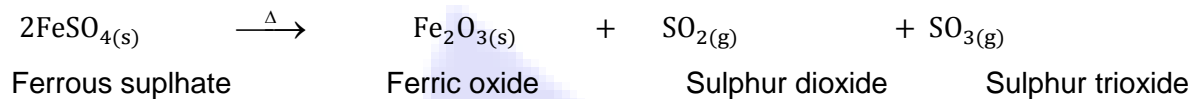


Combination reaction  $\text{A} + \text{B} \rightarrow \text{AB} + \text{Energy}$ .

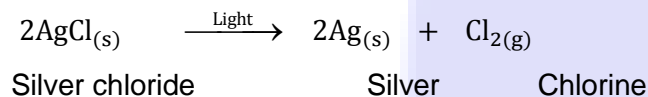


12. Write one equation each for decomposition reactions where energy is supplied in the form of heat, light or electricity.

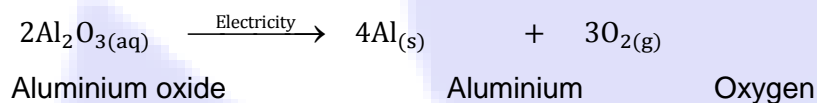
Ans. (a) Thermal decomposition.



(b) Decomposition by light.



(c) Decomposition by electricity.

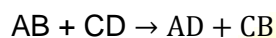


13. What is the difference between displacement and double displacement reactions? Write equations for these reactions.

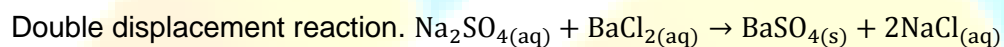
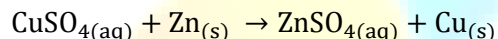
Ans. In a displacement reaction, a more reactive element replaces a less reactive element from a compound.



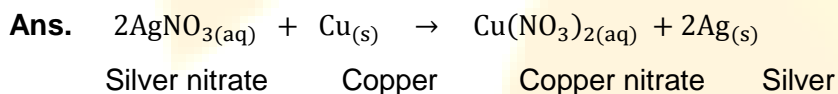
In a double displacement reaction, two atoms or a group of atoms switch places to form new compounds.



For example. Displacement reaction.



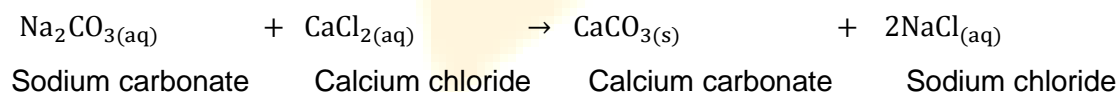
14. In the refining of silver, the recovery of silver from silver nitrate solution involved displacement by copper metal. Write down the reaction involved.



15. What do you mean by a precipitation reaction? Explain by giving examples.

Ans. A reaction in which an insoluble solid (called precipitate) is formed is called a precipitation reaction.

For example.



In this reaction, calcium carbonate is obtained as a precipitate. Hence, it is a precipitation reaction.

Another example of precipitation reaction is.



Sodium sulphate Barium chloride Barium sulphate Sodium chloride

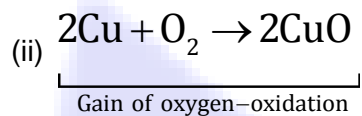
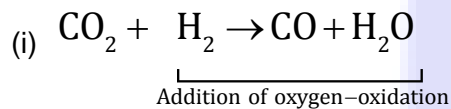
In this reaction, barium sulphate is obtained as a precipitate.

16. Explain the following in terms of gain or loss of oxygen with two examples each.

(a) Oxidation (b) Reduction

Ans. (a) Oxidation is the gain of oxygen.

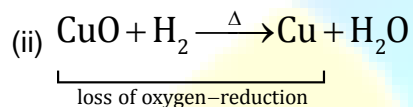
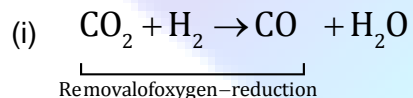
For example.



In equation (i), H<sub>2</sub> is oxidized to H<sub>2</sub>O and in equation (ii), Cu is oxidised to CuO.

(b) Reduction is the loss of oxygen.

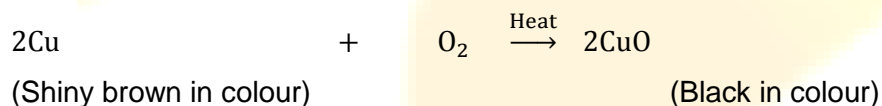
For example.



In equation (i), CO<sub>2</sub> is reduced to CO and in equation (ii), CuO is reduced to Cu.

17. A shiny brown-coloured element 'X' on heating in air becomes black in colour. Name the element 'X' and the black coloured compound formed.

Ans. 'X' is copper (Cu) and the black-coloured compound formed is copper oxide (CuO). The equation of the reaction involved on heating copper is given below.



18. Why do we apply paint on iron articles?

Ans. Iron articles are painted because it prevents them from rusting. When painted, the contact of iron articles from moisture and air is cut off. Hence, rusting is prevented. So presence of air and moisture is essential for rusting to take place.