

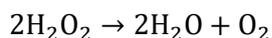
MOLE CONCEPT & STOICHIOMETRY

SOLVED EXAMPLE

1. "When stating the volume of a gas, the pressure and temperature should be also given." Why?
- Ans.** It is because, the volume of a gas changes, if the pressure or temperature or both change. Thus, while stating the volume of a gas pressure and volume has to be specified along with volume.
2. "The number of atoms in one mole of hydrogen is twice the number of atoms in 1 mole of helium at the same temperature and pressure." Why?
- Ans.** It is because; hydrogen is a diatomic gas, whereas helium is monoatomic gas. As number of atoms in one molecule of hydrogen is double, as compared to one molecule of helium, therefore one mole of hydrogen has double the atoms, as compared to helium at the same temperature and pressure.
3. Why is the term relative atomic mass used for atomic mass of an element?
- Ans.** Since the actual mass of an atom of element is extremely small, for comparing the masses of atoms of different elements, the mass of an atom of some light element is fixed as standard and the masses of other atoms are expressed relative to the standard mass, so the term relative atomic mass is used.
4. Why relative atomic mass is compared with $1/12$ of carbon?
- Ans.** Naturally occurring hydrogen has three isotopes (${}_1\text{H}^1$, ${}_1\text{H}^2$ and ${}_1\text{H}^3$) and its relative atomic mass works out 1.008 rather than 1. Atomic mass of carbon is exact 12. Thus, $1/12$ mass of carbon work out as 1. It is on account of the above reason $1/12$ mass of carbon is used for comparing relative atomic masses of other elements.
5. Why atomic weight of many elements is fractional?
- Ans.** The atomic weight of all the elements is never whole number but the atomic weights of many elements are fractional. The atomic weight of an element is fractional because the atomic weight of an element is actually the average relative weight of all the naturally occurring isotopes of the element.
6. What are the limitations of a chemical equation?
- Ans.** Following are the limitations of a chemical equation:
- (i) A chemical equation does not tell us the physical state of the reactants and the products in the reaction.
 - (ii) It does not tell us the actual concentration or dilution of the reactants used in the reaction.
 - (iii) It does not tell whether the reaction starts at its own or some heat is required to start the reaction.
 - (iv) It does not tell whether the reaction is violent in nature or not.
 - (v) The time taken by the reaction to complete itself is also not known.

7. What is the important information given by a balanced chemical equation of decomposition of hydrogen peroxide?

Ans. The decomposition of hydrogen peroxide is given by the equation:



The equation gives the following important information:

- (i) Hydrogen peroxide decomposes to form water and oxygen.
- (ii) Two molecules of hydrogen peroxide when decomposed form two molecules of water and one molecule of oxygen.
- (iii) 68 gm. of hydrogen peroxide when decomposed, yield 36 gm. of water and 32 gm. of oxygen.

8. A compound is formed by 24 grams of X and 64 grams of oxygen, if X = 12 and O = 16. Find the simplest formula of the compound."

Ans.

Relative number of atoms

$$X = \frac{24}{12} = 2$$

$$O = \frac{64}{16} = 4$$

Simplest ratio

$$X = \frac{2}{2} = 1$$

$$O = \frac{4}{2} = 2$$

Therefore, simple ratio between X and O is X : O = 1 : 2

Thus, empirical formula of the compound is XO_2 .

9. What are the applications of Avogadro's law?

- Ans.**
- (i) It determines the molecule formula of a gas.
 - (ii) It determines atomicity of gases.
 - (iii) It explains Gay-Lussac's law of combining volumes.
 - (iv) It establishes the relation between molecular weight and vapour density of a gas.
 - (v) It establishes relation between gram molecular weight and grass molecular air volume of a gas.
 - (vi) Avogadro's hypothesis modifies Dalton's atomic theory.

10. What do you understand by the statement "The vapour density of CO_2 is 22"? The molecular weight of the gas is twice its vapour density.

Ans. Molecular weight = 2 × molecular weight of carbon dioxide

$$\text{CO}_2 = 12 + 2(16) = 12 + 32 = 44 \text{ gram.}$$

The vapour density of carbon dioxide

$$\text{V.D. (CO}_2) = \frac{\text{Molecular weight of CO}_2}{2}$$

$$\text{V.D. (CO}_2) = \frac{44}{2}$$

$$\text{V.D. (CO}_2) = 22$$

The vapour density of CO_2 is 22 indicates that the mass of x litres of hydrogen gas or the molecular mass of CO_2 is 44 gm.

thus, the vapour density of carbon dioxide (CO_2) is 22.

11. How will you differentiate between atomic weight and actual weight of an element?

Ans.

Atomic weight	Actual weight
(i) It is the number of times an atom of an element is heavier than a carbon atom (C^{12}).	It is the weight of an atom as compared to a standard weight of 1g.
(ii) Its numerical value is quite large and is conveniently used for comparing the weights of various atoms and making calculations.	Its numerical value is quite small and is inconvenient to use in comparing the weights of various atoms and making calculations.
(iii) It represents the mass of 6.023×10^{23} atoms.	It represents the mass of one atom.
(iv) It is simply a number as it expresses the ratio between two weights.	It is the absolute weight of an atom and is generally expressed in grams.

12. Explain the following:

(i) Is it possible to change the temperature and pressure of a fixed mass of gas without changing its volume ?

(ii) One mole of hydrogen contains $2 \times 6.023 \times 10^{23}$ atoms of hydrogen whereas one mole of helium contains 6.023×10^{23} atoms of helium.

Ans. (i) Yes, an increase in temperature produces an increase in volume which can be reduced (changed) to original volume by the increase in pressure.

(ii) Hydrogen is a diatomic gas.

So one molecule of hydrogen = 2 atoms

\therefore 1 mole or 6.023×10^{23} molecules of $H_2 = 2 \times 6.023 \times 10^{23}$ atoms

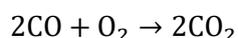
On the other hand, helium is monoatomic gas,

\therefore One molecule of helium 1 atom of He

or 6.023×10^{23} molecules of He = 6.023×10^{23} atoms of He.

13. How does Avogadro's law explain Gay-Lussac's law of gaseous volumes?

Ans. Avogadro's law states that equal volumes of all gases contain equal number of molecules under similar conditions of temperature and pressure. Since, when gases react chemically, they do so in volumes which bear a simple whole number ratio to each other and to the volume of products, provided the products are also in gaseous state under similar conditions of temperature and pressure. This is what Gay Lussac's law states. For example, in the reaction of carbon monoxide with oxygen, two volumes of carbon monoxide react with one volume of oxygen to give two volumes of carbon dioxide under similar conditions of temperature and pressure.



2 vol. 1 vol. 2 vol.

The volume ratio of carbon monoxide, oxygen and carbon dioxide is 2: 1: 2.

14. Under the same conditions of temperature and pressure, you collect 2 litres of carbon dioxide, 3 litres of chlorine, 5 litres of hydrogen, 4 litres of nitrogen and 1 litre of sulphur dioxide. In which gas will there be sample?

(i) The greatest number of molecules

(ii) The least number of molecules. Justify your answer.

Ans. The greatest number of molecules will be in 5 litres of hydrogen and the least number of molecules in 1 litre of sulphur dioxide. The justification is based on Avogadro's law, which states that equal volumes of all gases, under conditions of same temperature and pressure, contain same number of molecules. So greater the volume, greater will be the number of molecules.

15. What is the relationship between gram molecular weight and gram molecular volume at S.T.P.?

Ans. Density of gas is defined as mass per unit volume. Volume is usually taken as 1 dm³ at S.T.P.

$$\text{Density of gas} = \frac{\text{Mass of gas}}{1 \text{ dm}^3 \text{ of gas at S. T. P}}$$

Molar volume of hydrogen:

Density of hydrogen = 0.09 gm/dm³ at S.T.P.

Gram molecular weight of hydrogen = 2.016 g

0.09 g of hydrogen occupies volume = 1 dm³ at S.T.P.

2.016 g of hydrogen occupies volume = dm³ at S.T.P.

= 22.4 dm³ at S.T.P.

As 2.016 g of hydrogen = 1 gram molecular weight

∴ 1 gram molecular weight of hydrogen occupies 22.4 dm³ at S.T.P.

16. State Gay-Lussac's Law of combining volumes.

Ans. The law states that—Under same conditions of temperature and pressure, the volume of gases taking part in a chemical reaction show simple whole number ratio to one another and to those of products if gaseous.

17. (i) What is 'mole scale' of a compound?

(ii) What is the relation between atomic mass and equivalent mass?

Ans. (i) The molecular weight of a compound expressed in gram is known as "a mole" of a compound.

The multiples of the fractions of a mole give different mole values of a compound. The molecular weight of carbon dioxide is 44; hence 44 gm is one mole of carbon dioxide. 88 gm, and 176 gm of carbon dioxide represent 2 moles and 4 moles respectively.

In general,

$$\text{Number of moles of a compound} = \frac{\text{Mass of the compound in gram}}{\text{Molecular weight of the compound}}$$

(ii) Atomic mass = Equivalent mass × Valency.

18. Under the same conditions of temperature and pressure, one collects 2.2 liter of CO₂, 3.3 liters of Cl₂, 5.5 liters of hydrogen, 4.4 liters of nitrogen and 1.1 liters of SO₂. In which gas sample there will be:

(i) Greatest number of molecules. (ii) The least number of molecules.

Ans. We know that 22.4 liters of any gas at S.T.P. has 6.023×10^{23} molecules. If the volume of gas at S.T.P. is more than 22.4 liters, then the number of molecules will be greater and vice-versa.

(i) 5.5 liters of hydrogen will contain greatest number of molecules.

(ii) 1.1 liters of sulphur dioxide will contain least number of molecules.

19. (i) When gases react together, their reaction volume bears a simple ratio to each other, under the same conditions of temperature and pressure. Who proposed this gas law?

(ii) What is the volume (measured in dm³ or litres occupied by one mole of gas at S.T.P.?)

Ans. (i) This law was proposed by "Gay Lussac".

(ii) One mole of gas occupies 22.4 litres at S.T.P.