

WORKSHEET-1

- Write correct balanced equations for each of the following :
 - When potassium chlorate is strongly heated in a hard glass test tube.
 - When red lead [Pb_3O_4] is heated.
 - When carbon dioxide is passed over red hot coke.
 - When ammonium nitrate is heated
 - When steam is passed over red hot iron.
- Give two examples of industrial processes involving
 - oxidation
 - reduction.
- In the equations given below, state giving reasons, whether the substances printed in bold have been oxidised or reduced.
 $\text{PbO} + \text{CO} \rightarrow \text{Pb} + \text{CO}_2$
 $\text{H}_2\text{S} + \text{Cl}_2 \rightarrow 2\text{HCl} + \text{S}$
- Explain what would you observe and write chemical equations, when :
 - Copper (II) carbonate is heated?
 - Lead (II) nitrate is heated?
 - Washing soda is heated?
- Write balanced equations for :
 - Reaction between steam and red hot iron.
 - Action of heat on copper nitrate crystals.
 - Action of heat on ammonium chloride.

Which of the above reactions involves sublimation, decomposition and reversible reactions?
- Describe what happens or what would you observe, when each of the following substances are heated in separate test tubes.
 - Zinc carbonate
 - Lead dioxide.
- State what will be observed and write a chemical equation when lead nitrate is heated in a test tube?
- Give balanced equations, wherever possible, or where this is not possible, explain the following by means of examples :
 - A reaction which gives out heat.
 - A reaction which takes place with the help of sunlight.
 - A reaction which is brought about by electric current
- Name four factors which favour a chemical reaction. Support your answer by one example in each case.
- What do you understand by the term "Reversible reaction" ? Give two examples.