

WORKSHEET-1

1. Give four differences between true solution and colloidal solution.
2. Briefly describe how will you carry out the following separations :
 - (i) Ammonium chloride from a mixture of ammonium chloride and sand without heating the mixture.
 - (ii) Ammonium chloride from a mixture of ammonium chloride and sand on heating.
 - (iii) Iron filings from sulphur.
 - (iv) All the constituents of a mixture of iron, sulphur and sand.
 - (v) All the constituents of a mixture of saw dust, sulphur and common salt.
3. From the techniques (methods) of distillation, filtration , fractional distillation, chromatography, electrolysis, crystallisation select and write down the techniques you would use to separate :
 - (i) Constituents of colouring matter in ink
 - (ii) Hydrated copper sulphate from its aqueous solution.
 - (iii) Pure copper from a sample of impure copper
 - (iv) Nitrogen gas from liquid air.
 - (v) Unused zinc after reaction of excess of zinc with dilute sulphuric acid.
4. How will you separate the following :
 - (i) Black copper oxide from a mixture of copper oxide and zinc oxide?
 - (ii) Unused magnesium from a reaction between magnesium and hydrochloric acid?
 - (iii) Copper filings from a mixture of copper filings and iron filings.
5. Why does a white deposit appear near the top of a test tube in which ammonium chloride is heated?
6. How will you obtain pure sulphur from a mixture of coke and sulphur?
7. How will you obtain a pure sample of iodine from a mixture of iodine and ammonium chloride, without heating?
8. How will you obtain pure iodine from a sample of impure iodine?
9. You are given a powdered mixture of sulphur, potassium chloride and carbon. State briefly how you would separate and collect each constituent in the solid or powder form?
10. Amongst dilute sulphuric acid, carbon disulphide and water which liquid will you select to remove sulphur sticking to a crucible?